

# The Electronic Structure and Stability of Germanium Tubes $\text{Ge}_{30}\text{H}_{12}$ and $\text{Ge}_{33}\text{H}_{12}$

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## Abstract

Geometries of non-tetrahedral and ultrastable silicon and germanium nanocrystals  $X_{18}H_{12}$  and  $X_{19}H_{12}$  ( $X = Si, Ge$ ) have recently been predicted for the development of cluster-based nanomaterials for energy and microengineering purposes. To further explore the possibility of larger Ge clusters, we investigated in this work the molecular and electronic structure of **the** germanium tube  $Ge_{30}H_{12}$ , composed of six parallel, planar hexagons using DFT calculations. Insertion of Ge atoms at the center of three inner hexagons of  $Ge_{30}H_{12}$  leads to a  $Ge_{33}H_{12}$  tube, which is also an energy minimum structure. The electronic structure and molecular orbital shapes of these tubes can be predicted by the wavefunctions of a particle on a hollow cylinder model and a cylinder model. Different aromaticity indices including PDI,  $I_{ring}$ ,  $I_{NG}$ , MCI, and  $I_{NB}$ , as well as the electron localization function (ELF) were calculated to evaluate the electron delocalization and the aromaticity of the Ge tubes considered.

## 1. Introduction

Since the discovery of benzene in 1825 by Faraday, aromaticity was introduced as the key concept in the realm of organic chemistry to describe the stability, molecular structure and reactivity of many organic molecules.<sup>1-4</sup> Due to the persistent ambiguity in the definition of aromaticity and its non-observable nature, it has been a subject of continuing debate among chemists. Despite the controversy in the definition of aromaticity, aromaticity still remains as a useful, even fundamental, concept for the interpretation of structural and chemical properties of a variety of classes of organic molecules.<sup>5</sup> Moreover, the discovery of aromatic inorganic/organometallic clusters promises to expand its scope of application.<sup>6-11</sup> Many indices have been proposed in the literature for direct and quantitative measurement of aromaticity, which can be categorized into four main groups, including energy, structure, electron delocalization, and magnetic based indices.<sup>12,13</sup> Accordingly, it has been recommended that a set of indices, rather than a sole index, should be employed for characterization of aromatic compounds.<sup>14</sup> The final results would be more reliable if different aromaticity indices provided consistent results for a set of compounds. The necessity of different aromaticity indices has been discussed in several recent reviews.<sup>1, 5, 15-18</sup>

Triggered by benzene discovery, chemists have long been seeking to expand the concept of aromaticity to inorganic molecules, especially to analogous silicon compounds.<sup>19,20</sup> Regarding this interest, many attempts were made to synthesize and predict stable aromatic silicones. For example, Scheschkewitz *et al.* synthesized a silicon counterpart to benzene, with bulky organic substituents instead of hydrogen, showing high stability based on electron delocalization and aromaticity.<sup>21</sup>

Silicon and germanium nanocrystals and nanostructures have drawn the attention of researchers for their potential application in energy conversion, energy storage, light-emitting diode, and memory devices. Nowadays solar cells based on silicon nanocrystals find their way into the energy production market and germanium nanostructures are being studied as high-capacity anode materials for Li-ion batteries.<sup>22-26</sup>

Recently, Vach<sup>27</sup> predicted the aromaticity of the electron-deficient tubular  $\text{Si}_{19}\text{H}_{12}$ . The predicted  $\text{Si}_{19}\text{H}_{12}$  contains three parallel and planar hexagons and one additional Si atom located in the middle hexagon. The central Si atom is multicoordinated and Si-Si bonds are characterized by electron deficient properties. Electron deficiency keeps electrons delocalized through the whole structure, and ultimately provides aromatic character to the system. Vach's calculations showed that  $\text{Si}_{19}\text{H}_{12}$  is thermodynamically more stable than  $\text{Si}_{18}\text{H}_{12}$  and such a stability is related to a strong electron delocalization.<sup>27</sup> Using electronic and magnetic criteria, it was shown that the over-coordinated  $\text{Si}_{19}\text{H}_{12}$  has electron deficient bonds and is more aromatic than benzene. Due to electron-deficient aromaticity, electronic and optical properties of  $\text{Si}_{19}\text{H}_{12}$  are totally different from the  $\text{Si}_{18}\text{H}_{12}$  nanotube,<sup>28</sup> which may result in light-harvesting applications such as solar cells and optoelectronic devices.<sup>29</sup> It was also shown that  $\text{Ge}_{18}\text{H}_{12}$  nanocrystals, similar to  $\text{Si}_{18}\text{H}_{12}$ , could be further stabilized by insertion of one germanium atom into the center of the middle hexagon, and exhibit some electron-deficient aromaticity. Our previous results showed that the electron-deficient aromaticity concept could be extended to Ge based materials.<sup>30</sup> The presence of such aromaticity is intriguing and has stimulated us to further explore the possibility of larger Ge clusters being stabilized by this bonding phenomenon.<sup>31</sup>

It is likely that the feature observed in tubular silicon and germanium species arises from a type of aromaticity that is called tubular or cylindrical aromaticity, as recently reported by some

of authors of the present study.<sup>32, 33</sup> This aromaticity is explained by the hollow cylinder model (HCM) in which the Schrödinger equation can be solved for a particle moving in a hollow cylindrical box.<sup>34</sup> Accordingly, the shape of the eigenstates obtained from the HCM is quasi-identical with that of molecular orbitals calculated from quantum chemical methods.

In this context, we set out to examine the electronic structures of two extended tubular germanium tubes including Ge<sub>30</sub>H<sub>12</sub> and Ge<sub>33</sub>H<sub>12</sub> making use of the hollow cylinder model (HCM) and cylinder model (CM), respectively. Also, different electronic and magnetic indices are analyzed to further shed light on the aromaticity of these compounds. The proposed germanium tubes in this study and their silicon counterparts are expected to play a role in the following generation of energy storage and energy conversion devices.

## 2. Computational Methods

Standard electronic structure calculations and geometry optimizations are performed within the framework of density functional theory (DFT) using hybrid B3LYP functional and the 6-31G(d) basis set<sup>35-37</sup> with the aid of the Gaussian 09 program.<sup>38</sup> Harmonic vibrational frequency calculations at the same level reveal only real vibrational frequencies and thus reported structures are minima of the potential energy surface. As for aromaticity indices, several criteria based on electron delocalization measures are employed.<sup>12, 13, 39</sup> These indices measure the cyclic electron delocalization of mobile electrons in aromatic rings. First, we study the *para*-delocalization index (PDI)<sup>17,35</sup> which is obtained using the delocalization index (DI)<sup>41,42</sup> as defined in the framework of the QTAIM of Bader.<sup>43-45</sup> The PDI is an average of all DI of *para*-related atoms in a given six-

membered ring. For mono-determinantal closed-shell wavefunctions, the DI between atoms  $A$  and  $B$  is given by:

$$\delta(A, B) = 4 \sum_{i,j}^{occ.MO} S_{ij}(A)S_{ij}(B) \quad (1)$$

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The summations in Eq. (1) run over all occupied molecular orbitals (MOs).  $S_{ij}(A)$  is the overlap between MOs  $i$  and  $j$  within the basin of atom  $A$ .  $\delta(A, B)$  provides a quantitative indication on the number of electrons delocalized or shared between atoms  $A$  and  $B$ . If the atoms are defined in the Hilbert space, the latter quantity is known as the Mayer bond order.<sup>46</sup> We also employ a set of four multicenter indices, namely, the  $I_{ring}$ ,  $I_{NG}$ , MCI, and  $I_{NB}$ . For a ring structure represented by the  $A=[A_1, A_2, \dots, A_n]$  string, the multicenter index ( $I_{ring}$ ) of Giambiagi *et al.*<sup>15</sup> for a closed-shell mono-determinantal wavefunction is defined as follows:

$$I_{ring}(A) = 2^N \sum_{i_1, i_2, \dots, i_N}^{occ.MO} S_{i_1 i_2}(A_1) S_{i_2 i_3}(A_2) \cdots S_{i_N i_1}(A_N) \quad (2)$$

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Some of us proposed a normalized version of the  $I_{ring}$  index,<sup>47</sup>  $I_{NG}$ , which is expected to be less dependent on the ring size than its unnormalized analogs, and for aromatic species is given by:

$$I_{NG}(A) = \frac{\pi^2}{4NN_\pi} I_{ring}^{1/N} \quad (3)$$

$$\underline{I_{NG}(\mathcal{A}) = \frac{\pi^2}{4NN_\pi} I_{ring}^{1/N}} \quad (3)$$

where  $N$  is the total number of atoms in the ring and  $N_\pi$  the total number of  $\pi$  electrons.  $I_{NG}$  has the peculiarity of reproducing the so-called TREPE<sup>48</sup> values at the Hückel MO theory.<sup>49</sup>

According to Bultinck and coworkers<sup>50</sup>, summing up all the  $I_{ring}$  values resulting from the permutations of indices  $A_1, A_2, \dots, A_N$  defines a new index of aromaticity, the multicenter index (MCI) whose formula reads:

$$MCI(A) = \frac{1}{2N} \sum_{P(A)} I_{ring}(A) \quad (4)$$

$$\underline{MCI(\mathcal{A}) = \frac{1}{2N} \sum_{P(\mathcal{A})} I_{ring}(\mathcal{A})} \quad (4)$$

where  $P(A)$  stands for a permutation operator acting over string  $A$  interchanging the atomic labels  $A_1, A_2, \dots, A_N$  to generate up to the  $N!$  permutations of the elements in the string. Generally, the values of MCI and  $I_{ring}$  are in tight correlation because the dominant contribution to MCI is the Kekulé structure, nonetheless some exceptions may arise.<sup>47</sup> Finally, there is a normalized version of the MCI index for aromatic rings<sup>47</sup>,  $I_{NB}$ , given by:

$$I_{NB}(A) = \frac{C}{NN_\pi} [2N \cdot MCI(A)]^{1/N} \quad (5)$$

$$\underline{I_{NB}(\mathcal{A}) = \frac{C}{NN_\pi} [2N \cdot MCI(\mathcal{A})]^{1/N}} \quad (5)$$

where  $C \approx 1.5155$ . For the series of indices defined above, the higher the PDI,  $I_{ring}$ ,  $I_{NG}$ , MCI, and  $I_{NB}$  values, the more aromatic a six-membered ring is. The expressions  $I_{NB}$  and  $I_{NG}$  used in this work do not include the term  $G(N_\pi)$  which was given in the original article.<sup>47</sup> It is worth mentioning

that none of the multicenter indices aforescribed can be easily computed for rings of more than twelve members.<sup>51</sup>

Calculation of atomic overlap matrices (AOM) and computations of the DI,  $I_{ring}$ ,  $I_{NG}$ , MCI, and  $I_{NB}$  are performed with the AIMPAC<sup>52</sup> and ESI-3D<sup>53-55</sup> collection of programs. Calculations of the DIs with DFT cannot be performed exactly because the electron-pair density is not available at this level of theory.<sup>56</sup> As an approximation, we use the Kohn-Sham orbitals obtained from a DFT calculation to compute Hartree-Fock-like DIs through Eq. (4) and, therefore, we do not expect to recover electron correlation effects.

We also analyze the electron localization function (ELF) using the TopMod program.<sup>57</sup> As shown by Savin *et al.*,<sup>58</sup> the ELF measures the excess of kinetic energy density due to the Pauli repulsion. In the region of space where the Pauli repulsion is strong the ELF is close to one, whereas where the probability of finding same-spin electrons close together is high, the ELF tends to zero. For an  $N$ -electron single determinantal closed-shell wavefunction built from Hartree-Fock (HF) or Kohn-Sham orbitals, the ELF is given by<sup>59, 60</sup>:

$$\eta = \frac{1}{1 + \left(\frac{D}{D_h}\right)^2}, \quad (6)$$

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Where

$$D = \frac{1}{\rho} \nabla_2^2 \Gamma^{(2)\sigma\sigma}(\vec{r}_1, \vec{r}_2) = \frac{1}{2} \sum_{j=1}^N |\nabla \phi_j|^2 - \frac{1}{8} \frac{|\nabla \rho|^2}{\rho}, \quad (7)$$

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$$D_h = \frac{3}{10} (3\pi^2)^{2/3} \rho^{5/3}, \quad (8)$$

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$$\rho = \sum_{j=1}^N |\phi_j|^2, \quad (9)$$

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and where  $N$  is the number of electrons and  $\Gamma^{(2)\sigma\sigma}(\vec{r}_1, \vec{r}_2)$  is the same-spin contributions to the pair density. Since the ELF is a scalar function, an analysis of its gradient field can be carried out to locate its attractors (local maxima) and the corresponding basins. There are basically two chemical types of basins: the core basins (C) and the valence (V) ones, which are characterized by their synaptic order, i.e., the number of core basins with which they share a common boundary.<sup>61</sup> Graphical representations of the bonding are obtained by plotting isosurfaces of the ELF. These isosurfaces delimit volumes within which the Pauli repulsion is rather weak. The localization domains are called irreducible when they contain only one attractor, and reducible otherwise. The reduction of reducible domains is another criterion of discrimination between basins, and the reductions occur at a critical value of the bonding isosurface. The domains are ordered with respect to the ELF critical values, yielding bifurcations (tree diagrams). The ELF bifurcation values can be also taken as a measure of aromaticity. According to Santos and co-workers<sup>62</sup>, aromatic

compounds are characterized by high bifurcation values and small differences in bifurcation values of different basins.

The electronic structure of Ge<sub>30</sub>H<sub>12</sub> is also examined by the hollow cylinder model (HCM)<sup>32, 33</sup> while the electronic structure of Ge<sub>33</sub>H<sub>12</sub> is examined by the cylinder model (CM).<sup>63</sup> The latter can be considered as a special case of the former when the inner radius of the hollow cylinder is neglected. The Schrödinger equation for a particle moving in a hollow cylinder was solved by Gravesen and co-workers<sup>34</sup> and later on by Miliodoros for a cylinder with Möbius topology.<sup>64</sup> The key point, which causes a difference between the two models is related to the different boundary conditions, i.e., the CM is a special case of the HCM when the inner radius (R<sub>0</sub>) is neglected. A change in the boundary conditions leads to a change in the Schrödinger equation solution for P(ρ) function among the separable solution  $\psi(\rho, \theta, z) = P(\rho)\Theta(\theta)Z(z)$ . This leads to a change in the ordering of eigenstates (wavefunctions), such that the cylinder model becomes more suitable for a narrow nanotube-like molecular structure.<sup>63</sup>

In general, the energy spectrum in both HCM and CM is quantized with three quantum numbers, namely, the rational  $k$  ( $k = 1, 2, 3, \dots$ ), the rotational  $l$  ( $l = 0, \pm 1, \pm 2, \pm 3, \dots$ ), and the radial  $n$  ( $n = 1, 2, 3, \dots$ ), and it is defined by the expressions:

$$E = \frac{\hbar^2}{2mR^2} \left( (k_{ln}R)^2 + \left( \frac{k\pi}{L^f} \right)^2 \right) \quad (10)$$

$$E = \frac{\hbar^2}{2mR^2} \left( (\beta_{ln}R)^2 + \left( \frac{k\pi}{L^f} \right)^2 \right) \quad (11)$$

for HCM (eq.10) and CM (eq.11), respectively. In Eqs. (10) and (11),  $m$  is the mass of the particle moving in the cylinder,  $R$  is the radius of the cylinder,  $L^f = L/R$  with  $L$  being the length of the cylinder, and  $k_{ln}R$  is obtained from boundary conditions.

### 3. Results and Discussion

The optimized structures of  $\text{Ge}_{30}\text{H}_{12}$  and  $\text{Ge}_{33}\text{H}_{12}$  and the nomenclature used for Ge atoms are shown in Figure 1. The tube  $\text{Ge}_{30}\text{H}_{12}$  consists of five  $\text{Ge}_6$  rings containing 12 GeA, 12 GeB, and 6 GeC atoms, piled on top of each other. The tube  $\text{Ge}_{33}\text{H}_{12}$  is also comprised of two  $\text{Ge}_6$  rings (12 GeA) at the outermost and three innermost  $\text{Ge}_6$  units (12 GeB, 6 GeC, 2 GeE, and one GeD atoms). In hosting three additional Ge atoms along the symmetry axis, the tubular  $\text{Ge}_{30}\text{H}_{12}$  maintains a prism identity. The inserted GeE atom is positioned in the center of the structure, and 2 of GeD atoms are deviated from the center of the six neighbor Ge atoms by about 0.2 Å. By introducing Ge atoms inside the  $\text{Ge}_{30}\text{H}_{12}$  cage, the Ge–Ge bonds in three innermost and two outermost units are elongated by about 0.2 and 0.1 Å, respectively. No significant changes in the angles are found upon insertion of Ge atoms, except for the H-GeA-GeB angle values that change from 130° in  $\text{Ge}_{30}\text{H}_{12}$  to 123° in  $\text{Ge}_{33}\text{H}_{12}$ .

Our previous results pointed out that after introduction of one Ge atom into  $\text{Ge}_{18}\text{H}_{12}$ , the resulting  $\text{Ge}_{19}\text{H}_{12}$  structure stabilizes 4.5 eV with respect to the isolated units.<sup>30</sup> We attribute this stabilization to the increased aromatic character, as proved by different analysis. In this work, upon insertion of three Ge atoms into the  $\text{Ge}_{30}\text{H}_{12}$  tube,  $\text{Ge}_{33}\text{H}_{12}$  stabilizes 21.0 eV (see Figure 1). In the two following subsections, we analyze in detail the electronic structure of both tubular forms.

#### 3.1 $\text{Ge}_{30}\text{H}_{12}$

The tubular  $\text{Ge}_{30}\text{H}_{12}$  includes 54 Ge-Ge and 12 Ge-H bonds and it ~~is~~ also contains 132 valence electrons. According to the Lewis theory, this molecule fits perfectly with a structure of 132 valence electrons organized in 66 2c-2e bonds. However, the natural geometry of the tubular

moiety tends to create some delocalized electrons within the structure and decrease the number of localized electrons in the structure.

Figure 2.a displays the atom-in-molecule (AIM) topology maps, including 54 Ge-Ge and 12 Ge-H bond critical points (BCP, in green), 29 ring critical points (RCP, yellow points), and 4 cage critical bonds (CCP, in blue). While the Laplacian  $\nabla^2\rho_r$  at 54 Ge-Ge BCPs is negative, for the 12 Ge-H BCPs is positive, suggesting that the 54 Ge-Ge bonds behave as classical covalent bonds, and the 12 Ge-H are closed-shell bonds.<sup>65, 66</sup> The electron densities at Ge-H BCPs are also remarkable larger than those at Ge-Ge BCPs, suggesting a great “donor-acceptor” character of the Ge-H bonds.

The  $\nabla^2\rho_r$  Laplacian contour maps in different planes of  $\text{Ge}_{30}\text{H}_{12}$  are shown in Figure 3.a to 3.e. Figure 3.a displays great charge concentration regions belonging to the hydrogen atoms. Figure 3.c – 3.e show that the GeA-GeA, GeB-GeB, and GeC-GeC bonds display a similar electronic distribution, while Figures 3.a – 3b show small charge concentration in GeB-GeC bonds. Although the Laplacian  $\nabla^2\rho_r$  at the 54 Ge-Ge BCPs is negative, the Laplacian values at 12 GeB-GeC BCPs are close to zero, suggesting less localized electrons in GeB-GeC bonds than in the other Ge-Ge bonds. The Mayer bond orders<sup>46, 67</sup> of  $\text{Ge}_{30}\text{H}_{12}$  given in Table 1 also show larger values for the Ge-H bonds than for the GeB-GeC bonds. In order to understand these properties, we now examine the electronic structure of the tubular  $\text{Ge}_{30}\text{H}_{12}$  by analyzing the interaction between the MOs of  $\text{Ge}_{30}$  skeleton and the MOs of 2  $[\text{H}_6]$  units (2 x  $[\text{H}_6]$ ).

The electron structure of small nanotubes such as the series boron double ring  $\text{B}_{2n}$  ( $n = 10-14$ ),<sup>33</sup> the triple ring tubular  $\text{B}_{27}^+$ ,<sup>32</sup> and the skeleton  $\text{Si}_{15}$  in the triple ring  $\text{Mn}_2@\text{Si}_{15}$ <sup>68</sup> have been investigated using the hollow cylinder model (HCM).<sup>63</sup> In these anti-prism structures, valence MOs can be separated in three different sets: 2c-2e MOs (s-MOs), radial MOs (r-MOs) and

tangential (t-MOs). Recently, the HCM has been applied to a prism structure, the skeleton  $\text{Si}_{12}$  of  $\text{Cr@Si}_{12}$ , whose whole valence MOs fit in a unique HCM.<sup>63</sup> The fact that only one HCM is required for all valence MOs is due to the  $\text{sp}^3$  hybridization of silicon atoms. However, the latter destabilizes the  $\text{Si}_{12}$  skeleton.

Germanium is isovalent with silicon, and the tubular skeleton  $\text{Ge}_{30}$  consists of five  $\text{Ge}_6$  rings. Therefore, we also find the whole set of valence MOs of the tubular skeleton  $\text{Ge}_{30}$  obey only one HCM with the parameters  $\varepsilon = 0.77$  and  $L^f = 5.85$ . The radius and the height of the skeleton  $\text{Ge}_{30}$  amount to  $9.87 \text{ \AA}$  and  $2.44 \text{ \AA}$ , respectively. The active radius  $r$  ( $r = \varepsilon \cdot R$ ) of  $1.88 \text{ \AA}$ , which is deduced from the HCM, is shorter as compared to the value of  $2.11 \text{ \AA}$  of the Ge van der Waals radius.<sup>69</sup> The height of the hollow cylinder model  $L$  ( $L = L^f * R$ ) is  $14.26 \text{ \AA}$ . Compared to the real height of  $9.86 \text{ \AA}$  of the skeleton  $\text{Ge}_{30}$  (i.e., the distance between the top and bottom rings of the skeleton), the height of the hollow cylinder in  $\text{Ge}_{30}$ , in which electrons move, is larger by  $\Delta L = 2.20 \text{ \AA}$  ( $\Delta L = (14.26 - 9.86)/2$ ). The active radius of HCM for the skeleton  $\text{Ge}_{30}$  of  $\text{Ge}_{30}\text{H}_{12}$ , and that of HCM applied to the skeleton  $\text{Si}_{12}$  of  $\text{Cr@Si}_{12}$  are not significantly different from each other, while the extended length in  $\text{Ge}_{30}$  is obviously much longer than that in  $\text{Si}_{12}$  ( $\text{Ge}_{30}$  includes 5 rings while  $\text{Si}_{12}$  has only 2 rings).

As previously indicated for the  $\text{Si}_{12}$  skeleton,<sup>70</sup> the different behavior in the three sets s-MOs, r-MOs, and t-MOs and the interaction with atomic orbitals of Cr lead to a separation of these sets in the silicon double ring. Therefore, the different behavior in three sets s-MOs, r-MOs, and t-MOs of the skeleton  $\text{Ge}_{30}$  becomes obvious. However, interaction between the MOs of  $\text{Ge}_{30}$  skeleton and the MOs of  $2 \times [\text{H}_6]$  will result in a transfer of delocalized electrons above and below the  $\text{Ge}_{30}$  tube to the localized Ge-H bonds in  $\text{Ge}_{30}\text{H}_{12}$ .

Thus, we can only consider the two classical sets: the  $\pi$ -MOs set from r-MOs and the  $\sigma$ -MOs set from s-MOs and t-MOs. In  $\text{Ge}_{30}\text{H}_{12}$ , we have 6 MOs in  $\pi$ -MOs set and 60 MOs in  $\sigma$ -MOs set. Figure 5 shows the formation of the set of  $\pi$ -MOs of  $\text{Ge}_{30}\text{H}_{12}$  by contributions of MOs from 2 x  $[\text{H}_6]$  into the  $\pi$ -MOs set of the  $\text{Ge}_{30}$  skeleton. It is easily recognized that all these MOs are Ge–H anti-bonding orbitals. With 20  $\pi$ -electrons, the  $\text{Ge}_{30}$  skeleton follows the  $4N + 2M$  rule with  $N = 3$  and  $M = 4$ ,  $N$  and  $M$  being the number of degenerated and non-degenerated  $\pi$ -MOs, respectively. Let us define the difference between MOs energies as:

$$\Delta\text{MO} = E_{\text{MO}(\text{Ge}_{30}\text{H}_{12})} - E_{\text{MO}(\text{Ge}_{30})}, \quad (12)$$

where  $E_{\text{MO}(\text{Ge}_{30}\text{H}_{12})}$  and  $E_{\text{MO}(\text{Ge}_{30})}$  are the energies of the MOs in  $\text{Ge}_{30}\text{H}_{12}$  and  $\text{Ge}_{30}$ , respectively, which are assigned the same quantum number from the HCM. All  $\Delta\text{MO}$  values are positive for the  $\pi$ -MOs interaction.

In both  $(k \ 0 \ 2)$ -orbital and  $(k \pm 1 \ 2)$ -orbital, the  $\Delta\text{MO}$  values increase along with the rational quantum number  $k$ . Therefore, in  $\text{Ge}_{30}$ , the occupied MOs with high  $k$  rapidly become the unoccupied MOs in  $\text{Ge}_{30}\text{H}_{12}$ . As a result, 10  $\pi$ -MOs in  $\text{Ge}_{30}$  are formed while the  $\pi$ -MOs set of  $\text{Ge}_{30}\text{H}_{12}$  has now 6 MOs. Hence, reduction of delocalized radial electrons is a way to stabilize the tubular structure.

In  $\text{Ge}_{30}\text{H}_{12}$ , there are 27  $\sigma$ -MOs with negative  $\Delta\text{MO}$  values. The 15 MOs having the largest  $|\Delta\text{MO}|$  values are shown in Figure 6. These tangential orbitals follow the  $4N + 2M$  rule with  $N = 4$  and  $M = 7$ . We can easily recognize that the large H contributions correspond to interactions between H orbitals and the  $p_z$ -orbitals of the GeA moieties, whereas the small H contributions correspond to interactions between H orbitals and s-orbitals of Ge atoms.

The results obtained for different indices of aromaticity are given in Table 2. In order to provide some reference values, let us mention that PDI,  $I_{\text{ring}}$ ,  $I_{\text{NG}}$ , MCI, and  $I_{\text{NB}}$  for benzene at the B3LYP/6-311G++(d,p) level are 0.103 e, 0.048 e, 0.041 e, 0.072 e, and 0.041 e, respectively.<sup>5</sup> The comparison of these values with those of Table 2, reveals that the aromaticity of the  $\text{Ge}_6$  rings is not particularly large. All indices considered indicate that the aromaticity follows the order  $\text{Ge}_6(\text{B}) > \text{Ge}_6(\text{C}) > \text{Ge}_6(\text{A})$ . The aromatic character of the most aromatic  $\text{Ge}_6(\text{B})$  ring within  $\text{Ge}_{30}\text{H}_{12}$  ranges from about 50% of that of benzene according to  $I_{\text{NG}}$  and  $I_{\text{NB}}$  indices to less than the 10% given by  $I_{\text{ring}}$  and MCI. Not unexpectedly, given the approximate  $\text{sp}^3$  hybridization of its Ge atoms, the least aromatic ring is the outer one.

To further analyze the aromaticity of  $\text{Ge}_{30}\text{H}_{12}$  we have also performed an ELF bifurcation analysis of this system (see Figure 7). We find that the separation of the six tangential Ge-Ge bond basins of the inner  $\text{Ge}_6$  rings ( $\text{Ge}_6(\text{B})$  &  $\text{Ge}_6(\text{C})$ ) occurs at a value  $\text{ELF} = 0.48$ . At  $\text{ELF} = 0.49$ , the Ge-Ge tangential bonding basins of the three inner  $\text{Ge}_6$  rings are separated from the radial valence Ge basins. At  $\text{ELF} = 0.51$ , separation of the Ge-Ge tangential bonding basins of the outer-most  $\text{Ge}_6$  ring (2 rings) takes place. Finally, at  $\text{ELF} = 0.58$ , separation of the radial valence Ge-Ge basins into two parts occurs and, at  $\text{ELF} = 0.60$ , a complete separation of radial basins is observed. According to these results, two conclusions can be drawn: i) First, the radial aromaticity ( $\text{BV}(\text{ELF})=0.60$ ) is higher than the tangential one ( $\text{BV}(\text{ELF}) = 0.51$ ), and ii) the tangential aromaticity of the different  $\text{Ge}_6$  rings is quite similar ( $\text{BV}(\text{ELF})$  ranging from 0.48 to 0.51).

### 3.2 $\text{Ge}_{33}\text{H}_{12}$

Figure 2.b shows the AIM maps of the  $\text{Ge}_{33}\text{H}_{12}$  electron density, including 74 Ge-Ge and 12 Ge-H BCPs (in green), 56 RCPs (yellow points), and 14 CCPs (in blue).

Although the Laplacian  $\nabla^2\rho_r$  values at the GeA-GeA, GeA-GeB, and GeA-H BCPs (which are masked by presence of GeE and GeD atoms) are not remarkably changed with respect to those of Ge<sub>30</sub>H<sub>12</sub>, the Laplacian  $\nabla^2\rho_r$  at the GeB-GeB, GeC-GeC, and GeB-GeC BCPs presents positive values, therefore, GeB-GeB, GeC-GeC, and GeB-GeC bonds become more closed-shell-like bonds.

The  $\nabla^2\rho_r$  Laplacian contour maps in different planes of Ge<sub>33</sub>H<sub>12</sub> are shown in Figures 3.f to 3.j. The latter Figures show a large charge concentration area in GeA-GeA, GeA-GeB, GeA-H, and GeE-GeD bonds. While Figures 3.f - 3.g indicate no charge concentration between the GeB-GeC bonds, Figures 3.i - 3.j display a reduced charge concentration area between the GeB-GeB and GeC-GeC bonds. Finally, Figure 3.f points out a large charge concentration area within the GeE-GeD bonds.

The presence of three Ge atoms inside the Ge<sub>33</sub>H<sub>12</sub> tube permits the circulation of electrons through the cylinder more freely than in the hollow Ge<sub>30</sub>H<sub>12</sub>. In our previous study,<sup>63</sup> the ground state valence MO of the triple ring B<sub>27</sub><sup>+</sup> and the fullerene-like B<sub>14</sub> have been used as a standard model to decide whether the hollow cylinder model or the cylinder model should be applied. Figure 4 displays a ground valence MO, HOMO – 40 of the Ge<sub>30</sub> skeleton (Figure 4.b) and a ground valence MO, HOMO – 45, of the Ge<sub>33</sub> skeleton (Figure 4.d). The shape of the HOMO – 40 of the Ge<sub>30</sub> skeleton suggests that the hollow cylinder model is suitable for the Ge<sub>30</sub> skeleton, whereas the shape of HOMO – 45 of the Ge<sub>33</sub> skeleton indicates that the cylinder model should be used for the latter compound.

Figure 8 shows a few MOs from the Ge<sub>33</sub> skeleton and the corresponding ones obtained from the cylinder model, proving the excellent predictability of the cylinder model for this kind of geometrical structures. In particular the shape of the (1 0 3) eigenstate, which corresponds to the

HOMO – 2 of the Ge<sub>33</sub> skeleton, it is not only reproduced but its relative energy is also correctly predicted by the cylinder model.

The parameters of the cylinder model derived for the Ge<sub>33</sub> skeleton are shown in Figure 4.c. The height of the cylinder model amounts to 12.34 Å, while the height of the Ge<sub>33</sub> skeleton is 10.30 Å. Accordingly, the height of the current cylinder model is extended by a small length of  $\Delta L = 1.02$  Å in each direction. The active radius of the cylinder model is 4.16 Å, while the average radius of the Ge<sub>33</sub> skeleton (from five rings) is 2.52 Å. This means that the electrons within the Ge<sub>33</sub> skeleton can move in a cylinder whose radius is larger than the radius of the structure by  $\Delta R = 1.62$  Å.

Although the Ge<sub>33</sub> skeleton is actually larger than the Ge<sub>30</sub> counterpart, the height of the cylinder in which the electrons of Ge<sub>33</sub> are moving becomes smaller than the height of the hollow cylinder in which electrons of Ge<sub>30</sub> circulate. The extended height of the cylinder of Ge<sub>33</sub> is also smaller than the extended height of the hollow cylinder of Ge<sub>30</sub>. The shape of the ground state valence MOs can partly be predicted considering this phenomenon.

According to the HCM, the electrons in the Ge<sub>30</sub> can move in and out of the tubular structure to distances up to 1.88 Å whereas, in Ge<sub>33</sub>, the CM predicts that the electrons can move in and out of the tubular structure up to 1.62 Å.

In general, the presence of three central Ge atoms vertically placed inside the tubular structure reduces the available volume for moving electrons. As it has three more Ge atoms, more valence electrons are added to the system, and the tubular Ge<sub>30</sub> also donates electrons to these central atoms. The net charges given in Table 3 show that large negative charges are located on GeE and GeD positions. Moreover, Table 1 displays that the largest Mayer bond orders correspond

to the GeE-GeD bonds, while the Mayer bond orders of GeB-GeB, GeB-GeC, GeC-GeC, GeB-GeD, and GeC-GeE bonds are significantly smaller.

Figure 9 shows the formation of the set of  $\pi$ -MOs of  $\text{Ge}_{33}\text{H}_{12}$  from contributions of MOs from 2 x  $[\text{H}_6]$  into the  $\pi$ -MOs set of  $\text{Ge}_{33}$  skeleton. At the first glance, the ground state  $\pi$ -MOs set is not formed by the  $p_{xy}$  hybrid, but rather by the  $s$  hybrid. However, formation of the  $(k\ 0\ 2)$ -orbitals still leads to anti-bonding orbitals / interactions between the  $\text{Ge}_{33}$  skeleton and hydrogen atoms 2 x  $[\text{H}_6]$ . The  $(4\ 0\ 2)$ -orbital and  $(5\ 0\ 2)$ -orbital are occupied MOs in the  $\text{Ge}_{33}$  skeleton, whereas, in  $\text{Ge}_{33}\text{H}_{12}$ , these orbitals are unoccupied.

The  $p_{xy}$  orbitals have larger contributions to the  $(k\ \pm 1\ 2)$ -orbitals. The  $(4\ \pm 1\ 2)$ -orbitals are strong bonding MOs with an energy difference of 2.24 eV between the HOMO – 13, HOMO – 13' of  $\text{Ge}_{33}\text{H}_{12}$  and HOMO – 3, HOMO – 3' of  $\text{Ge}_{33}$  skeleton. The  $(5\ \pm 1\ 2)$ -eigenstates of  $\text{Ge}_{33}\text{H}_{12}$  cylinder are assigned to the HOMO – 2 and HOMO – 3' with an energy of 0.38 eV lower than the HOMO and HOMO' of the  $\text{Ge}_{33}$  skeleton.

Figure 10 indicates formation of some  $\sigma$ -MOs with large  $\Delta\text{MO}$  values. Bonding MOs are found for the  $(k\ 3\ 1)$ -eigenstates, and  $(k\ \pm 1\ 1)$ -eigenstates with large  $k$  values. As a result, the delocalized basins located above and below the  $\text{Ge}_{33}$  skeleton become localized basins in the Ge-H bonds.

$\text{Ge}_{33}\text{H}_{12}$  consists of five  $\text{Ge}_6$  rings piled on top of each other. The outermost rings 1- $\text{Ge}_6$  and 5- $\text{Ge}_6$  (GeA atoms in Figure 1b), which are equivalent by symmetry, and the three innermost ones can also be considered as three  $\text{Ge}_7$  units (GeB and GeC atoms Figure 1b). Results in Table 4 show that the aromatic character of  $\text{Ge}_6$  rings is small for the outer rings, but substantial for the inner rings. All indices indicate that the aromaticity follows the ordering  $\text{Ge}_6(\text{B}) > \text{Ge}_6(\text{C}) > \text{Ge}_6(\text{A})$ . The aromatic character of the most aromatic  $\text{Ge}_6(\text{B})$  ring in  $\text{Ge}_{33}\text{H}_{12}$  ranges from about 50% (PDI,

$I_{NG}$ , and  $I_{NB}$ ) to less than the 5% given by  $I_{ring}$  and MCI. Overall, the outer rings of  $Ge_{33}H_{12}$ , as those of  $Ge_{30}H_{12}$ , have a low (if any) aromatic character. Conversely, the inner rings of  $Ge_{33}H_{12}$  have a significant aromatic character.

The ELF plots of  $Ge_{33}H_{12}$  (Figure 11) show that a separation of the six tangential Ge-Ge bonding basins of the three inner  $Ge_6$  rings ( $Ge_6(B)$  &  $Ge_6(C)$ ) occurs at  $ELF = 0.48$ . At this ELF value, there are Ge-Ge radial bonding basins between the three Ge atoms in the center of the cylinder. At  $ELF = 0.51$ , a separation of the Ge-Ge tangential bonding basins of the outermost  $Ge_6$  ring (2 rings) is observed. Then, at  $ELF = 0.58$ , the Ge-Ge radial bonding basins of the three innermost  $Ge_6$  units are separated from each other. At  $ELF = 0.60$ , the Ge valence basins in the tangential  $Ge_6(C)$  unit are separated from the radial Ge-Ge bonding basins, and, finally, at  $ELF = 0.62$ , the radial Ge-Ge bonding basins are completely separated from each other. From these ELF results, three conclusions emerge: i) First, the radial aromaticity ( $BV(ELF) = 0.62$ ) is larger than the tangential one ( $BV(ELF) = 0.51$ ), and ii) the tangential aromaticity of different  $Ge_6$  rings is quite similar to each other ( $BV(ELF)$  ranging from 0.48 to 0.51), 3) suggesting a similar aromatic character of both clusters  $Ge_{30}H_{12}$  and  $Ge_{33}H_{12}$ .

#### 4. Concluding Remarks

This paper investigates the geometries, chemical bonding, MOs analysis, and aromaticity of two hydrogenated germanium tubes  $Ge_{30}H_{12}$  and  $Ge_{33}H_{12}$ . We have shown that the MOs of  $Ge_{30}H_{12}$ , obtained from DFT calculations, are effectively reproduced by the interaction between the MOs of  $Ge_{30}$  skeleton, predicted by HCM, and the MOs of 12 hydrogen atoms. Also, the MOs of the  $Ge_{33}H_{12}$  structure can basically be reproduced using the cylinder model, considering  $Ge_{33}H_{12}$

as resulting from interaction of a  $\text{Ge}_{33}$  tubular skeleton and 12 hydrogen atoms. The aromaticity analysis of  $\text{Ge}_{30}\text{H}_{12}$  and  $\text{Ge}_{33}\text{H}_{12}$  suggests that both structures have similar aromatic character. In the two clusters, the inner rings show more aromatic character than the outer rings. Moreover, the radial aromaticity is found to be larger than the tangential counterpart. The results obtained showing the critical role of cylindrical aromaticity in the stability of these clusters could be important in potential emerging technologies.

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**Table 1.** Mayer bond orders of  $\text{Ge}_{30}\text{H}_{12}$  and  $\text{Ge}_{33}\text{H}_{12}$ .

<b>Cluster</b>	<b>Ge<sub>A</sub>-H</b>	<b>Ge<sub>A</sub>-Ge<sub>A</sub></b>	<b>Ge<sub>A</sub>-Ge<sub>B</sub></b>	<b>Ge<sub>B</sub>-Ge<sub>B</sub></b>	<b>Ge<sub>B</sub>-Ge<sub>C</sub></b>	<b>Ge<sub>C</sub>-Ge<sub>C</sub></b>	<b>Ge<sub>B</sub>-Ge<sub>D</sub></b>	<b>Ge<sub>C</sub>-Ge<sub>E</sub></b>	<b>Ge<sub>E</sub>-Ge<sub>D</sub></b>
<b>Ge<sub>30</sub>H<sub>12</sub></b>	0.886	0.872	0.760	0.801	0.688	0.743	—	—	—
<b>Ge<sub>33</sub>H<sub>12</sub></b>	0.892	0.855	0.734	0.554	0.486	0.567	0.508	0.436	1.087

**Table 2.** PDI (electrons),  $I_{ring}$  (electrons),  $I_{NG}$ , MCI (electrons), and  $I_{NB}$  results for the  $Ge_3H_{12}$  at the B3LYP/6-31g(d) level.

<b>Aromaticity index</b>	<b>Ge<sub>6</sub>(A)</b>	<b>Ge<sub>6</sub>(B)</b>	<b>Ge<sub>6</sub>(C)</b>
PDI	0.0066	0.0197	0.0134
$I_{ring}$	0.0004	0.0018	0.0010
$I_{NG}$	0.0183	0.0240	0.0216
MCI	0.0006	0.0026	0.0017
$I_{NB}$	0.0187	0.0237	0.0220

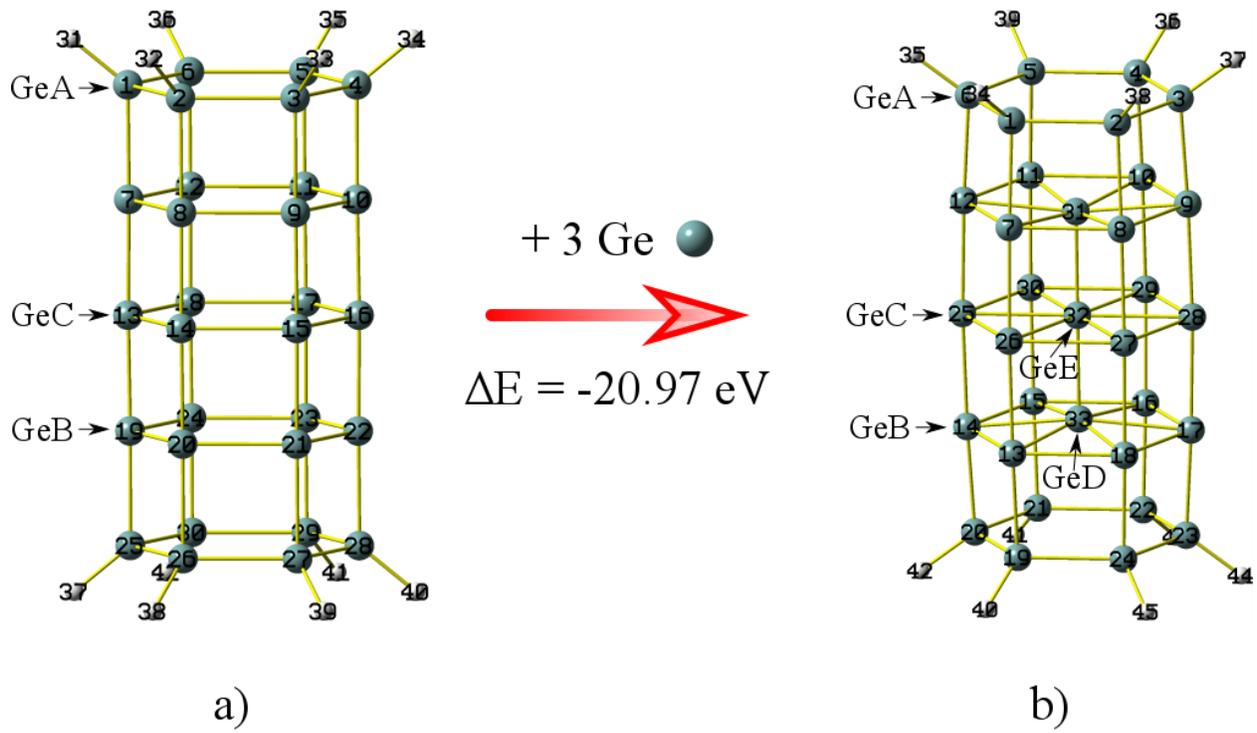
**Table 3.** The net charges (in electrons) of H and Ge atoms of Ge<sub>30</sub>H<sub>12</sub>.

<b>Cluster</b>	<b><math>q_H</math></b>	<b><math>q_{GeA}</math></b>	<b><math>q_{GeB}</math></b>	<b><math>q_{GeC}</math></b>	<b><math>q_{Ge-ins}</math></b>
<b>Ge<sub>30</sub>H<sub>12</sub></b>	-0.080	0.135	-0.084	0.056	—
<b>Ge<sub>33</sub>H<sub>12</sub></b>	-0.089	0.125	0.017	0.104	GeD: -0.332 GeE: -0.592

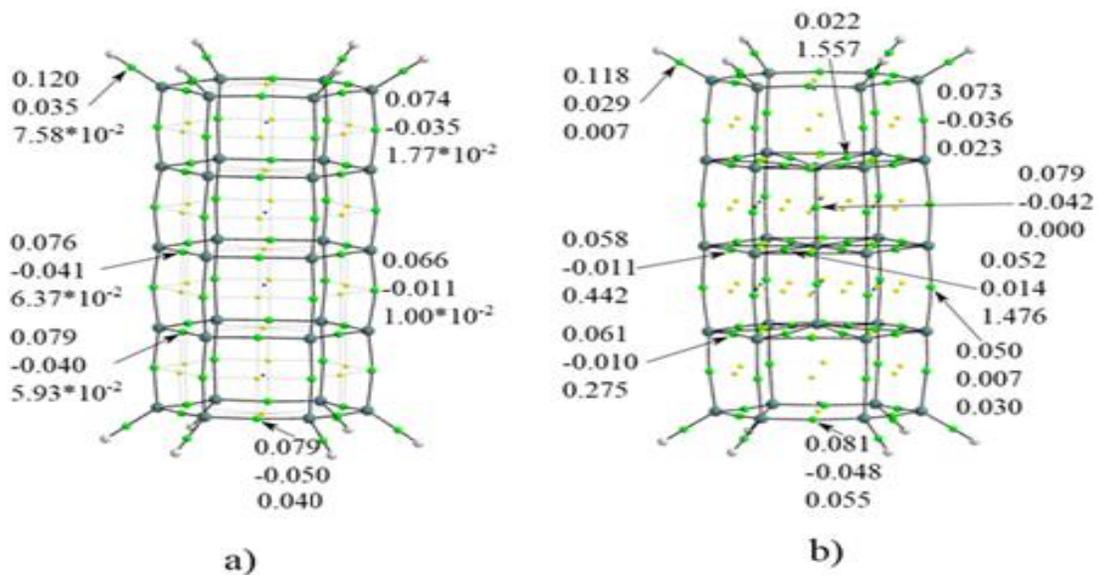
**Table 4.** PDI (electrons),  $I_{ring}$  (electrons),  $I_{NG}$ , MCI (electrons), and  $I_{NB}$  results for the  $Ge_{33}H_{12}$  at the B3LYP/6-31g(d) level.<sup>a</sup>

<b>Aromaticity index</b>	<b>Ge<sub>6</sub>(A)</b>	<b>Ge<sub>6</sub>(B)</b>	<b>Ge<sub>6</sub>(C)</b>
PDI	0.0053	0.0491	0.0329
$I_{ring}$	0.0003	0.0009	0.0007
$I_{NG}$	0.0173	0.0213	0.0203
MCI	0.0005	0.0027	0.0042
$I_{NB}$	0.0181	0.0237	0.0255

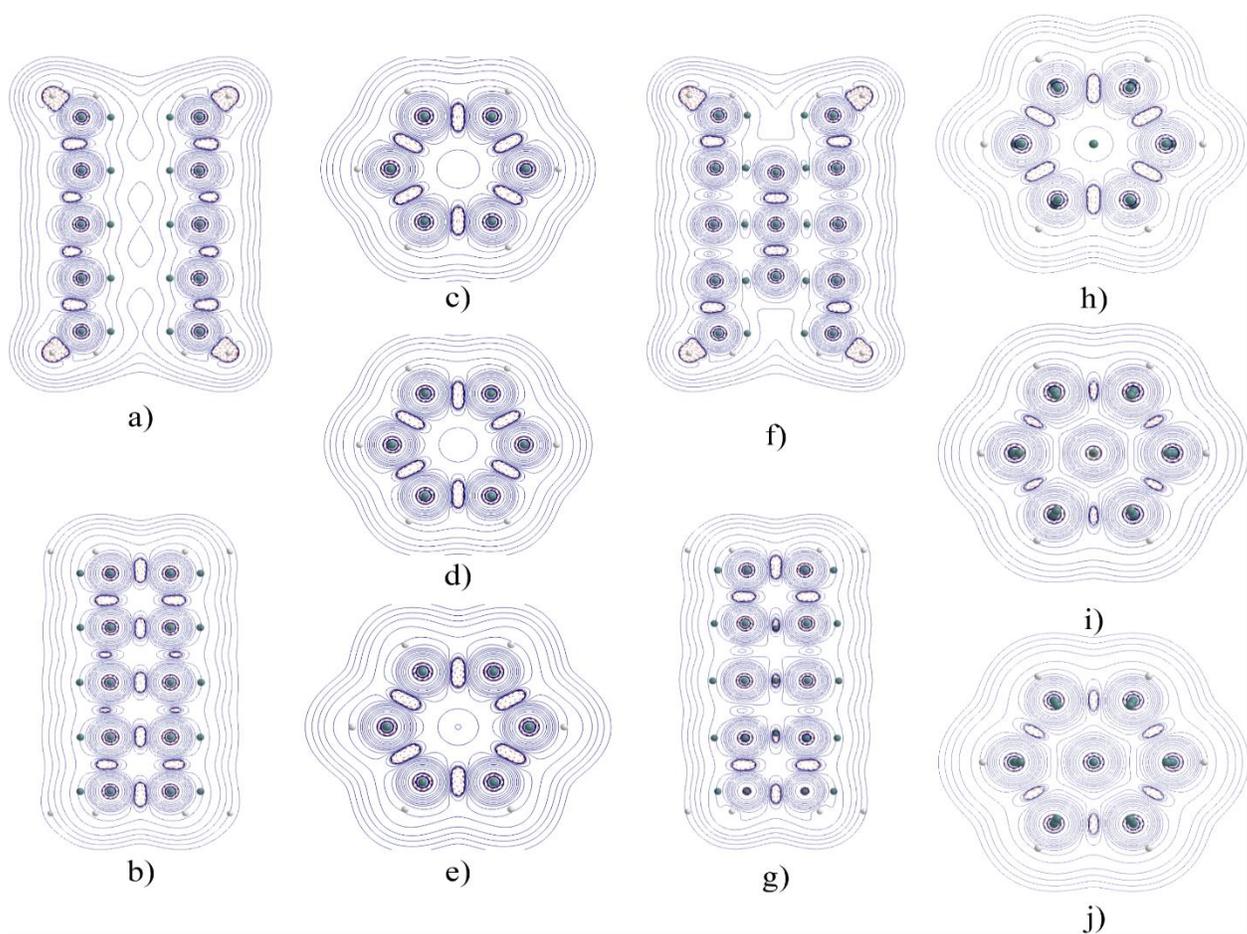
<sup>a</sup> MCI values of the seven-membered rings of Ge give MCI=-0.0011 for  $Ge_7$ (B) and MCI=0.0019 for 3- $Ge_7$ (C).



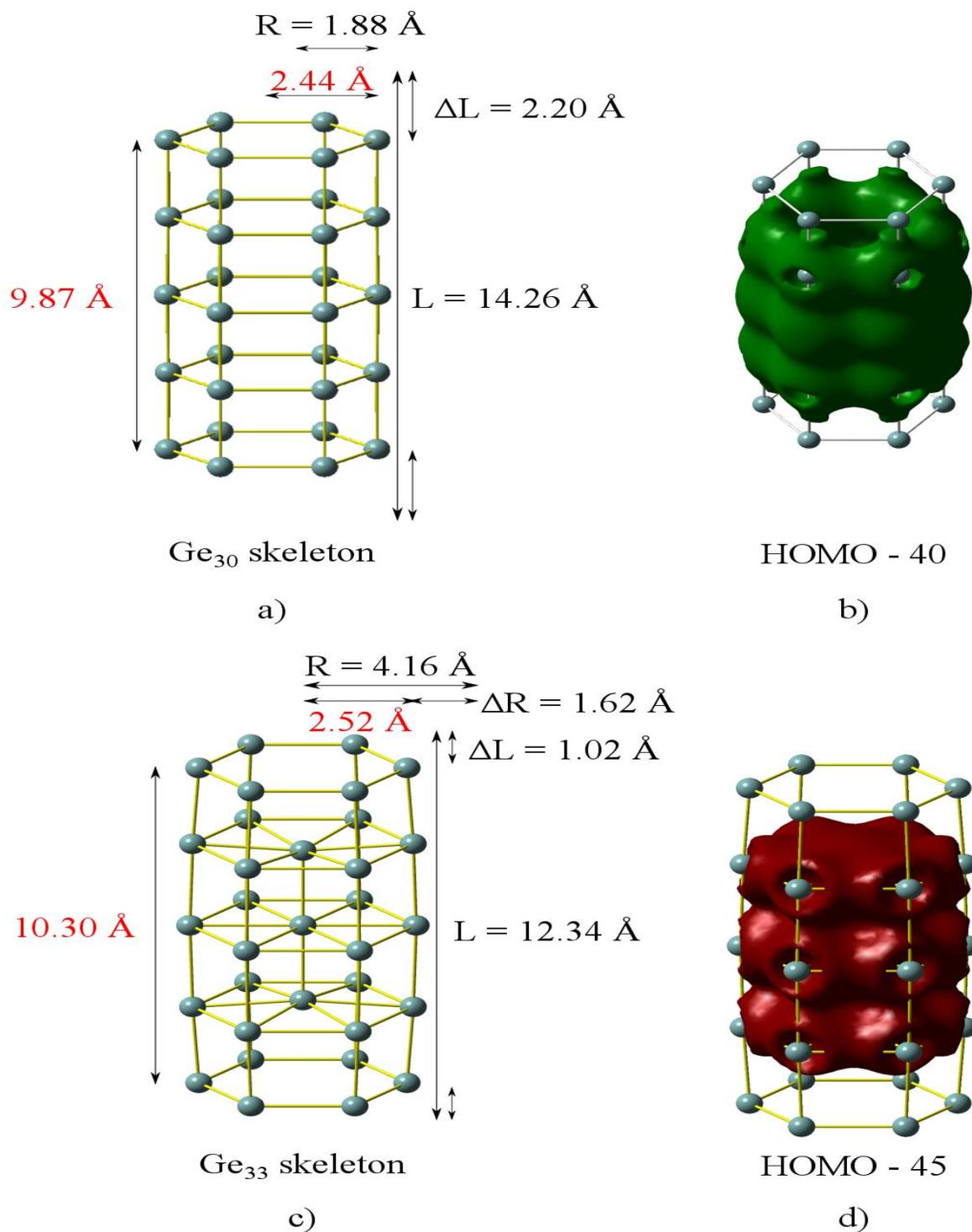
**Figure 1.** Optimized structures of a)  $\text{Ge}_{30}\text{H}_{12}$  and b)  $\text{Ge}_{33}\text{H}_{12}$  NCs along with the nomenclature used for Ge atoms.



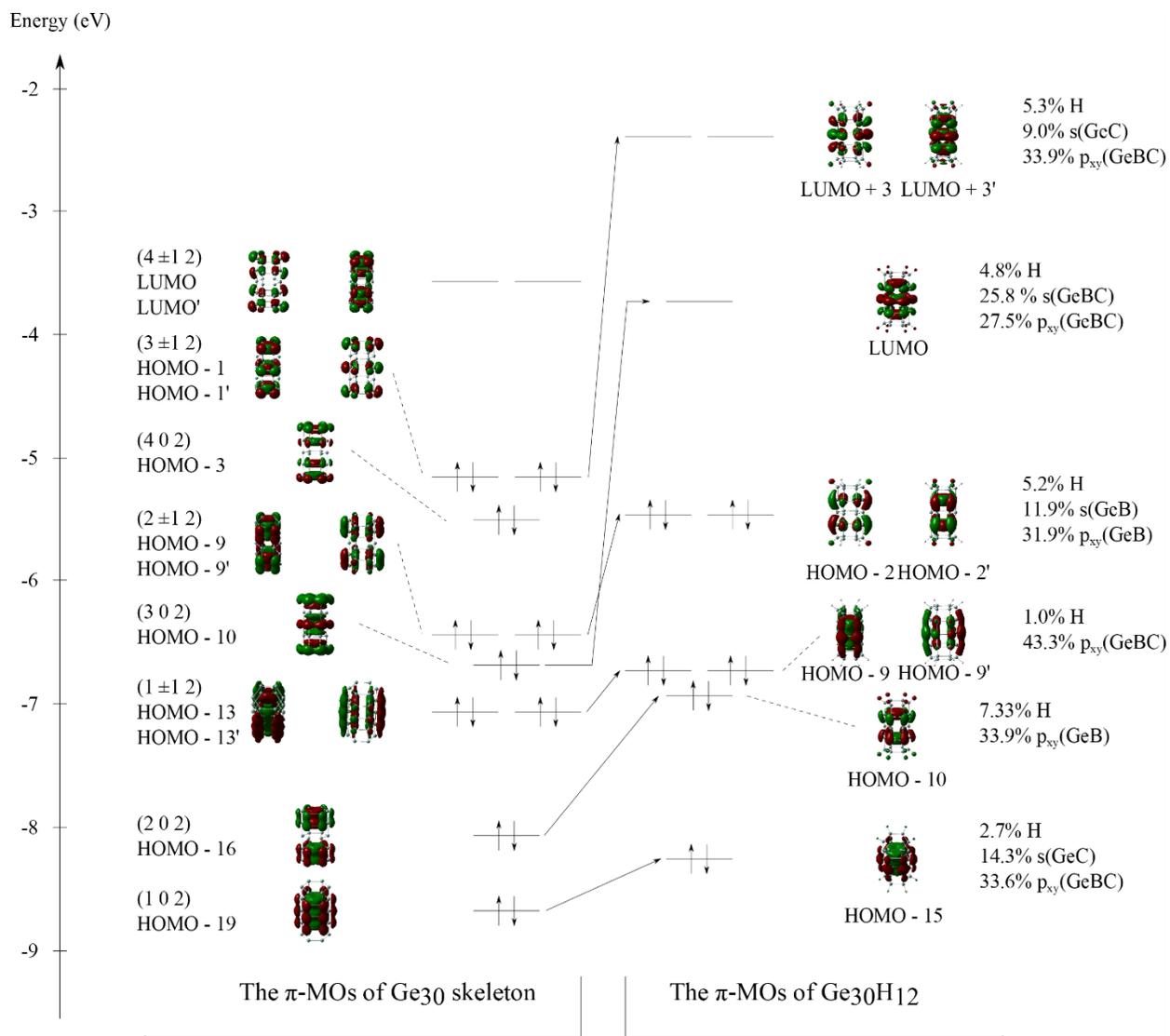
**Figure 2.** a) The AIM graphs of Ge<sub>30</sub>H<sub>12</sub> including 54 Ge-Ge and 12 Ge-H bond critical points (green points), 29 ring critical points (yellow points), and 4 cage critical bonds (blue points). b) The AIM graphs of Ge<sub>33</sub>H<sub>12</sub> including 74 Ge-Ge and 12 Ge-H bond critical points (green points), 56 ring critical points (yellow points), and 14 cage critical bonds (blue points). The value of electron density, Laplacian, and bond ellipticity at bond critical points are shown in three entries, one on top of each other.



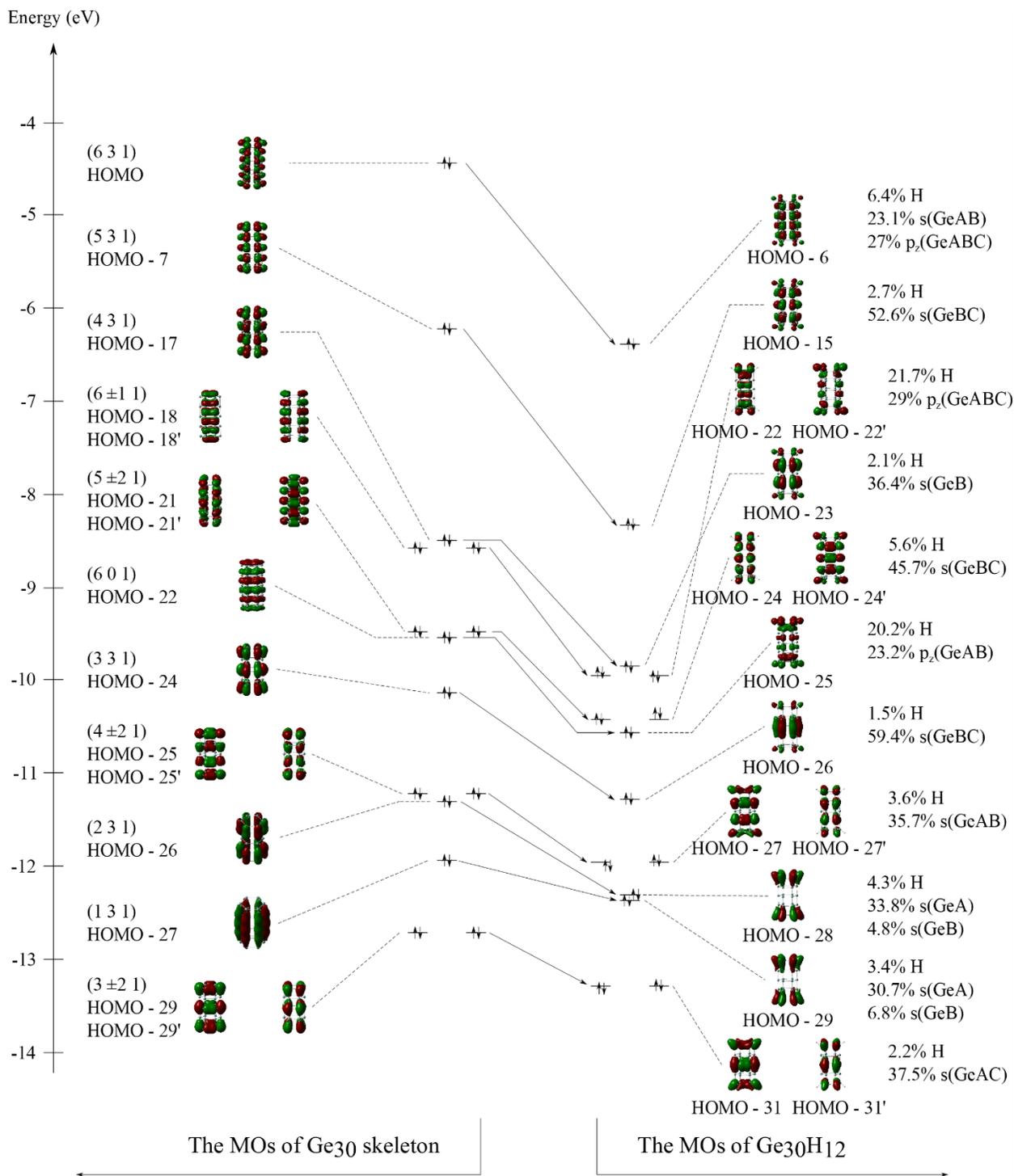
**Figure 3.** The Laplacian contour map for the  $\text{Ge}_{30}\text{H}_{12}$  in the plane: a) ( $\text{Ge}_1\text{Ge}_4\text{H}_{34}$ ), b) ( $\text{Ge}_8\text{Ge}_9\text{Ge}_{15}$ ), c) ( $\text{Ge}_1\text{Ge}_2\text{Ge}_3$ ), d) ( $\text{Ge}_7\text{Ge}_8\text{Ge}_9$ ), and e) ( $\text{Ge}_{13}\text{Ge}_{14}\text{Ge}_{15}$ ); and for  $\text{Ge}_{33}\text{H}_{12}$  in the plane: f) ( $\text{Ge}_3\text{Ge}_{23}\text{H}_{42}$ ), g) ( $\text{Ge}_1\text{Ge}_2\text{Ge}_{24}$ ), h) ( $\text{Ge}_1\text{Ge}_2\text{Ge}_3$ ), i) ( $\text{Ge}_7\text{Ge}_8\text{Ge}_9$ ), and j) ( $\text{Ge}_{26}\text{Ge}_{27}\text{Ge}_{28}$ ). Dashed lines indicate charge concentrations, solid lines indicate charge depletions. Labels of atoms shown in Figure 1.



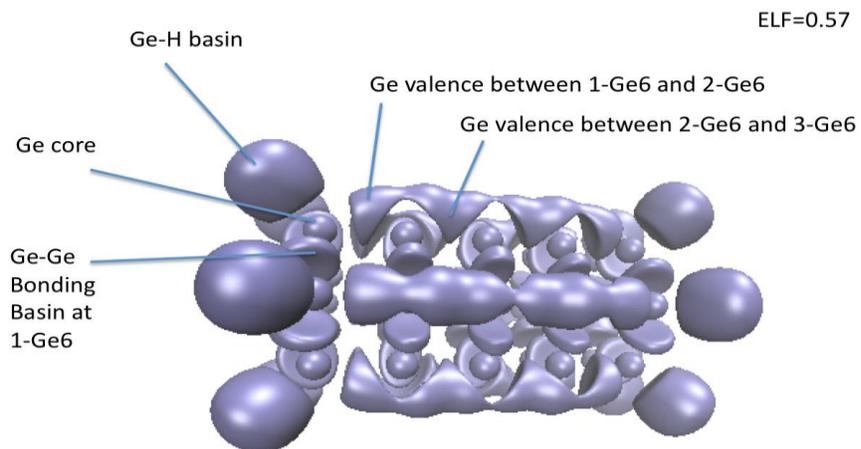
**Figure 4.** a) The Ge<sub>30</sub> skeleton of the Ge<sub>30</sub>H<sub>12</sub>. The height of the skeleton is 9.87 Å, while the height of the hollow cylinder in which the electron is moving in is 14.26 Å. The extended length is 2.20 Å. The average radius of the skeleton is 2.44 Å and the radius active of the hollow cylinder is 1.88 Å. b) The HOMO – 40 of the skeleton Ge<sub>30</sub>. c) The skeleton Ge<sub>33</sub> of the Ge<sub>33</sub>H<sub>12</sub>. The height of the skeleton is 10.30 Å while the height of the cylinder in which electron is moving in is 12.34 Å. The extended length is 1.02 Å. The average radius of the skeleton is 2.52 Å and the radius of the cylinder is 4.16 Å. d) The HOMO – 45 of the skeleton Ge<sub>33</sub>.



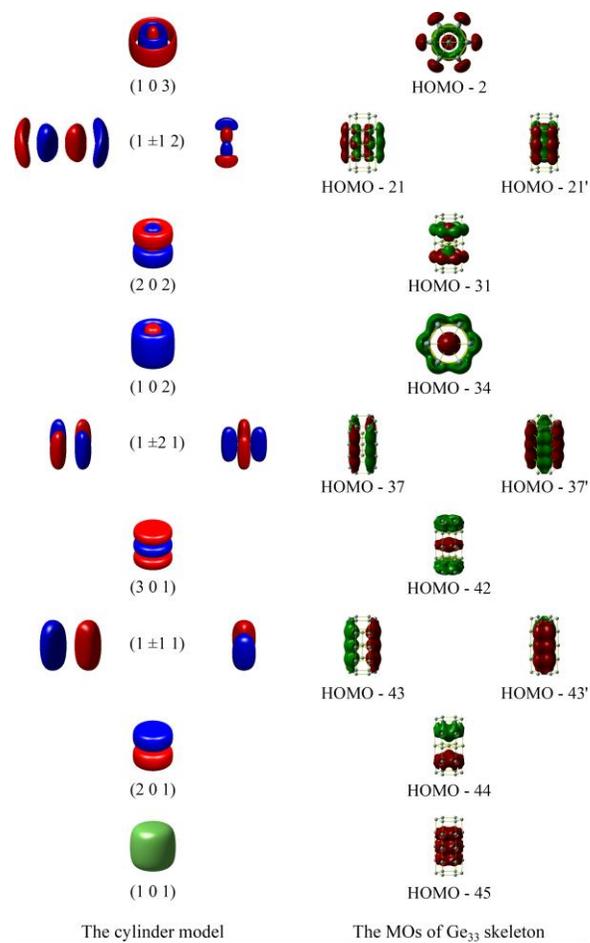
**Figure 5.** The formation of  $\pi$ -MOs set of  $\text{Ge}_{30}\text{H}_{12}$  from the interaction between  $\pi$ -MOs set of skeleton  $\text{Ge}_{30}$  and 12 hydrogen atoms.



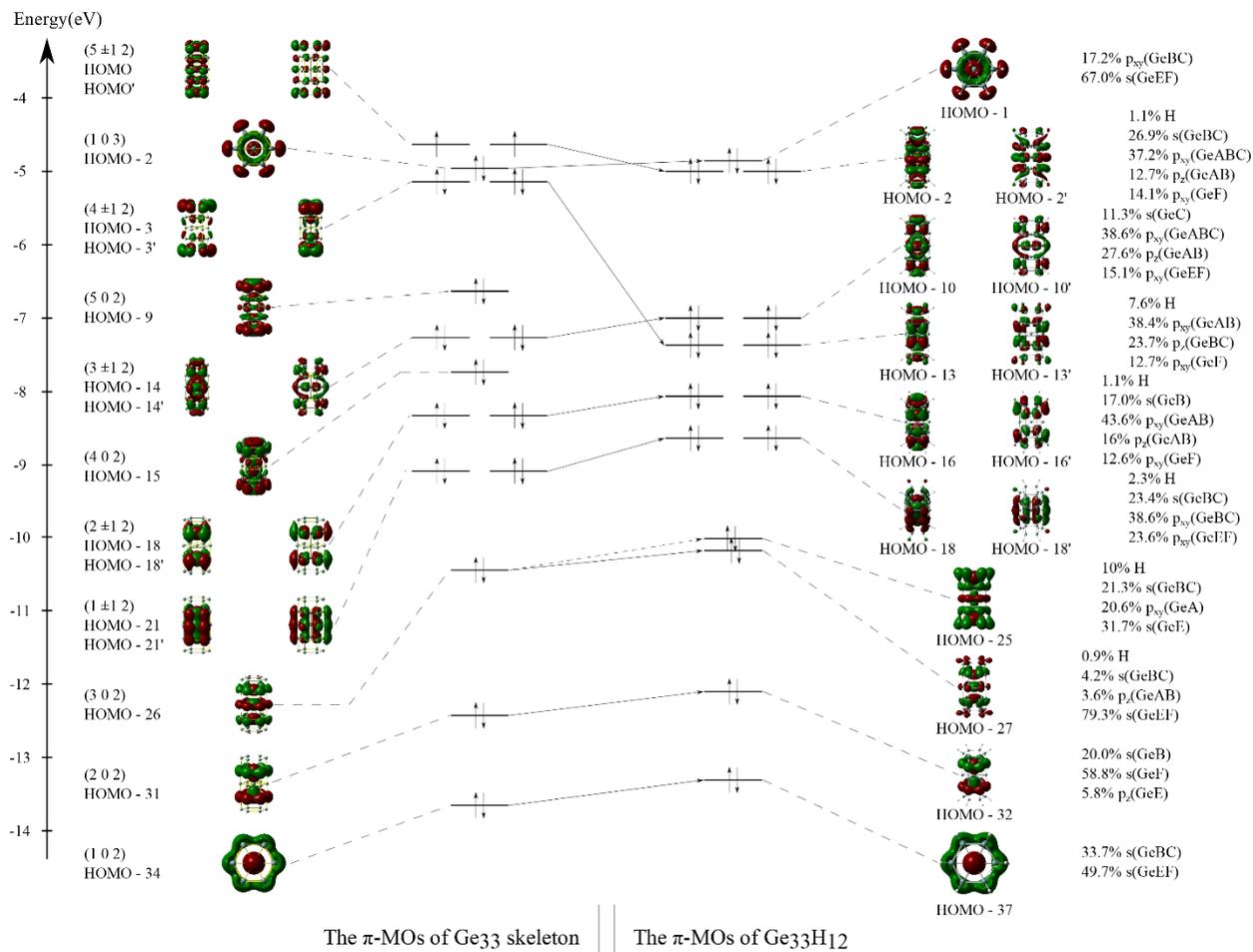
**Figure 6.** The formation of some  $\sigma$ -MOs of Ge<sub>30</sub>H<sub>12</sub> from the interaction between  $\sigma$ -MOs of skeleton Ge<sub>30</sub> and 12 hydrogen atoms which having the largest  $|\Delta MO|$  values.



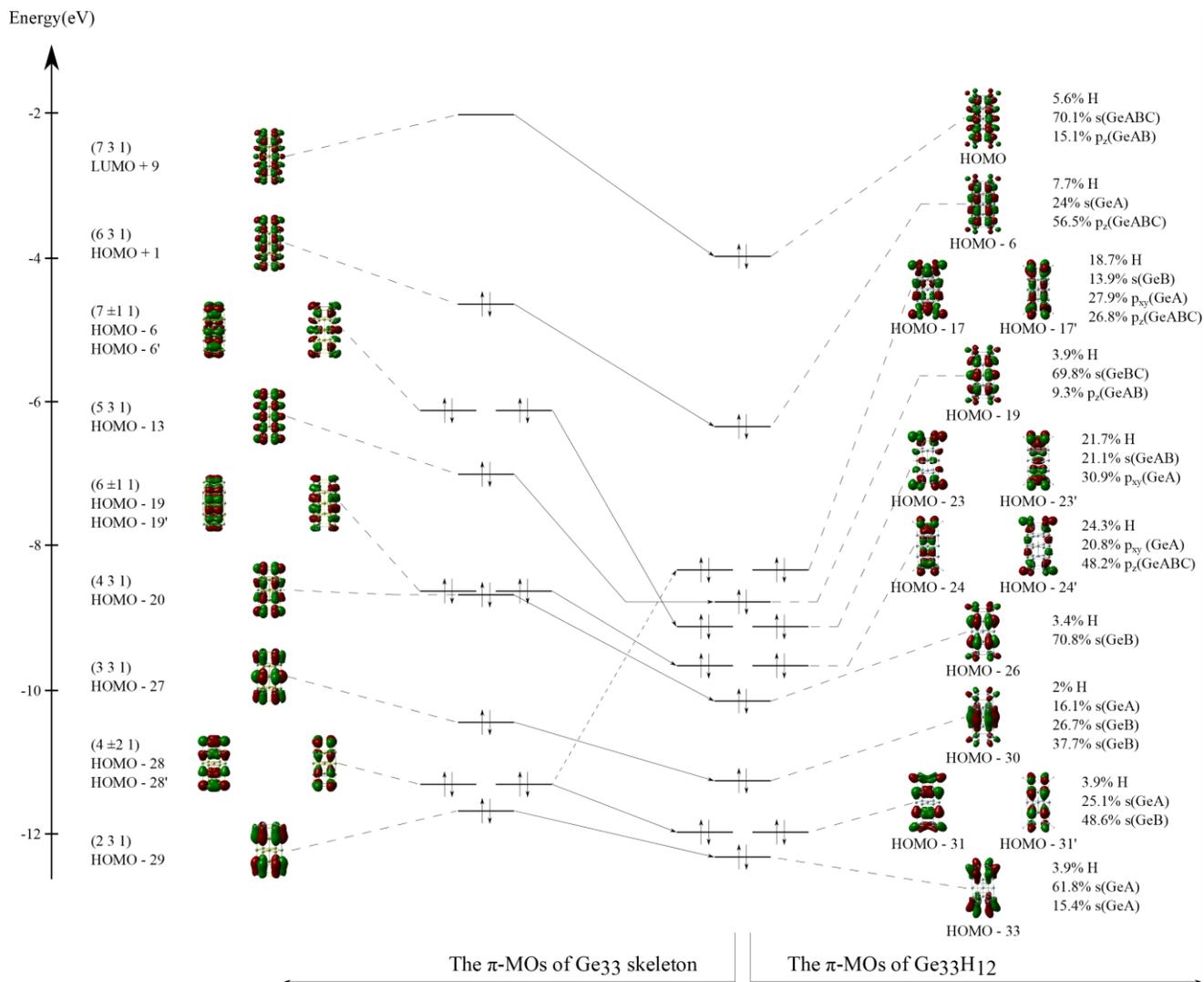
**Figure 7.** Electron localization function of Ge<sub>30</sub>H<sub>12</sub> obtained at the B3LYP/6-31G(d) level of theory. Isosurface value of 0.57.



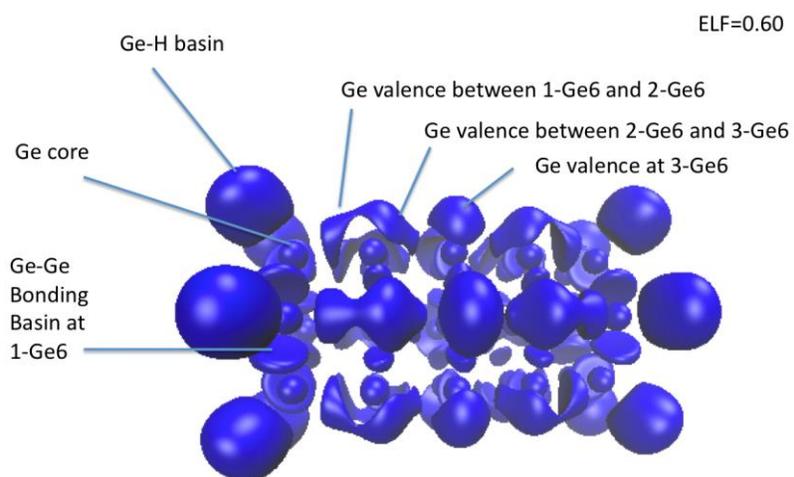
**Figure 8.** The comparison between the shape from cylinder model and the selected MOs from Ge<sub>33</sub> skeleton.



**Figure 9.** The formation of  $\pi$ -MOs set of  $\text{Ge}_{33}\text{H}_{12}$  from the interaction between  $\pi$ -MOs set of skeleton  $\text{Ge}_{33}$  and 12 hydrogen atoms,  $\text{Ge}_{33}$  skeleton are calculated at single point after removing 12H (without optimization).



**Figure 10.** The formation of some  $\sigma$ -MOs of Ge<sub>33</sub>H<sub>12</sub>, which, from the interaction between  $\sigma$ -MOs of skeleton Ge<sub>33</sub> and 12 hydrogen atoms.



**Figure 11.** Electron localization function of  $\text{Ge}_{33}\text{H}_{12}$  obtained at the B3LYP/6-31G(d) level of theory. Isosurface value of 0.60.

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